**Aim:**

**To investigate acid and metal reactions**

**Hypothesis:**

**if I have variety of metals i will get a variety of reactions**

**Method:**

1. **Get six test tubes. Fill to 5ml with acid. All must be level.**
2. **Place the metal into one test tube each.**
3. **observe results**
4. **place the test tubes in order of most reactive to least reactive**

**Discussion:**

1. **What gas was being produced in each reaction?**

**They released hydrogen or H2**

1. **Did you observe any difference between the reactions of metal in the acid?**

**Some of the reactions like copper didn’t even begin to react but magnesium burned up within seconds the rest sort of varied between bubbling a lot or having only a few bubbles**

1. **From Q2 write a list from the most reactive to the least reactive metal?**

**Mg, Zn, Pb, Fe, Al, Cu**

1. **If metals can be ranked in terms of their reactivity. What implications might this have on their usefulness?**

**They can be used for things which require them to last E.g. marine areas and boats need to have metals that can last. Iron is used a lot because it is strong and it last a long time. Things like magnesium would corrode and Copper although it was the strongest would cost too much and is weak.**